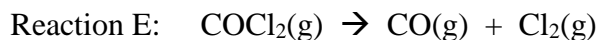
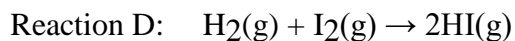
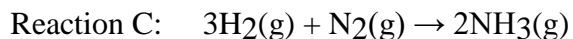
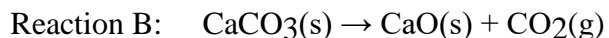
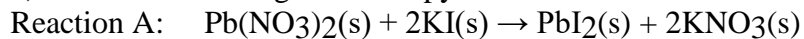


Crystal Ball questions on Explaining Entropy change

All of the following questions have not (as yet!) appeared in the NCEA Level 3 Exams

1) Describe the changes in entropy to the four reactions below and state reaction occurs with the largest increase in entropy change and why?



2) Discuss why the entropy of a gas is more than the entropy of the same substance in a solid state.

3) Write the equation for the dissolving of ammonium chloride in water and suggest how and why the entropy will change.

4) **extension:** Explain why a reaction with a negative value of ΔH and a negative value of ΔS is spontaneous only at low temperatures.